

# Effect of pH on Optical Properties of Graphene Oxide Quantum Dots

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**ABSTRACT**— Luminescent graphene oxide quantum dots (GOQDs) have attracted tremendous attention from scientists in chemistry, materials, biology, and physics science. They have specific properties such as low cytotoxicity, excellent electrochemical and optical properties, resistance to photo-bleaching, and good stability. In this study, GOQDs were synthesized using a simple and straightforward methodology. The synthesized GOQDs were characterized by Fourier Transform Infrared (FTIR) analyzer, ultraviolet–visible spectrophotometry (UV–VIS) absorption, Photoluminescence (PL) spectroscopy, and transmission electron microscopy (TEM) analyses. Then, optical properties of GOQDs such as absorption and luminescence with various pH values were investigated. GOQDs show absorption in the ultraviolet (UV) region and their position of photoluminescence peak is independent of pH value. The average size of QDs is less than 5 nm, as revealed by TEM. The GOQDs show green luminescence under UV irradiation (360 nm).

**KEYWORDS:** Graphene Oxide, Quantum Dots, pH, Photoluminescence, UV-Vis.

## 1. INTRODUCTION

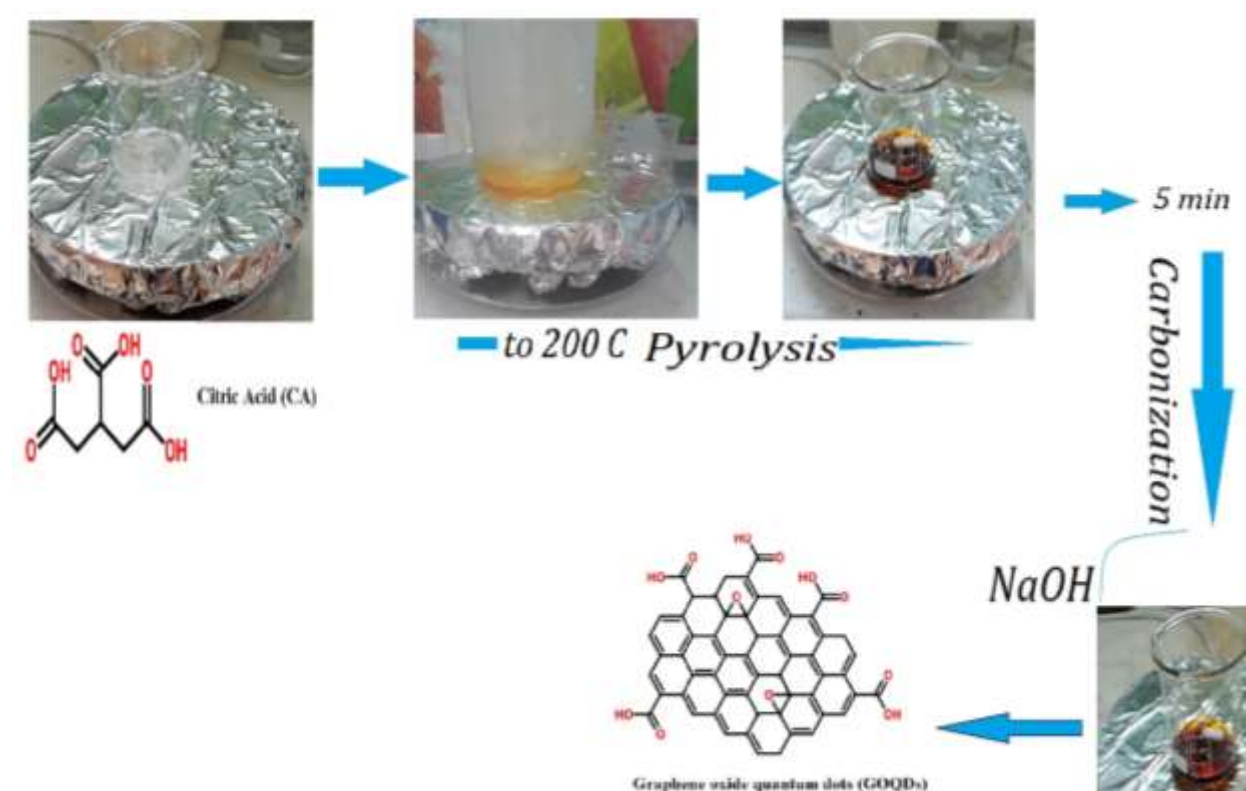
Quantum dots are called semiconductor nanocrystals with size of about 1 to 10 nm [1-3], which are widely used in light-emitting diodes (LEDs), lasers, environmental displays, electronic displays, and solar cells due to their unique electro-photo properties depending on

their size [4-7]. However, due to the high cost and toxicity of quantum dots based on, their use in industry has been limited [6]. In recent decades, as a cost-effective and promising alternative, graphene quantum dots (GQDs), graphene oxide quantum dots (GOQDs), and carbon quantum dots (CQDs) have emerged as a new class of nanomaterials [8, 9]. Graphene oxide (GO) is a modern, single-layer structure of the well-known material of graphite oxide. Graphite oxide is not a substance found in nature and is artificially produced as a single layer of carbon atoms with a sp<sup>2</sup> bond in a honeycomb network [10-12]. By reducing the particle size of graphene oxides to a multi-nanometer scale, the optical properties and hydrophobicity of graphene oxide can be adjusted. These particles are called “graphene oxide quantum dots.” Due to their very low toxicity, hydrophobicity, and high optical efficiency, GOQDs promise extensive medicine applications [13-15].

So far, many methods have been reported to synthesis of GO and GOQDs, that are usually requiring a long reaction time. Graphene derivatives are generally produced via two methods: “bottom-up” and “top-down”. The first method uses small molecular precursors such as citrate, carbohydrates, glucosamine, ascorbic acid, saccharides, via combustion or, solvothermal process to make graphene derivatives [16-21]. The second method (top-

down) usually breaks down more abundant carbon sources by means like electrochemical oxidation, laser ablation, and arc discharge [22-28]. Moreover, this should be mentioned that the above methods have risk of direct exposure to flammable carbon sources, strong oxidants and, strong acids, which may explode

or burn, the synthesis process is quite tedious and time-consuming. In this research, we will report a secure and easy one-pot synthesis method using citric acid (CA) as the starting material for synthesis GOQDs. This method is quite simple and does not require tedious and Time processes.



dispersed in ultrapure water. The synthesis process is shown in Fig. 1.

### C. Characterization

Characterizations of the synthesized GOQDs were conducted using TEM, FT-IR, UV-Vis, and, PL analyses. The TEM image of the GOQDs was acquired on a JEOL, JEM-2100F, 200 KV electron source. Fourier Transform Infrared (FTIR) analyzer spectrum in region 200–4000  $\text{cm}^{-1}$  were restored on an Equinox 55 FTIR spectrometer using Potassium bromide (KBr) pellets. Absorption spectra were recorded in area 200–900 nm by using Perkin Elmer 2500 UV-Vis spectrophotometer. The PL spectra were measured with G9800A Agilent, under ambient conditions.

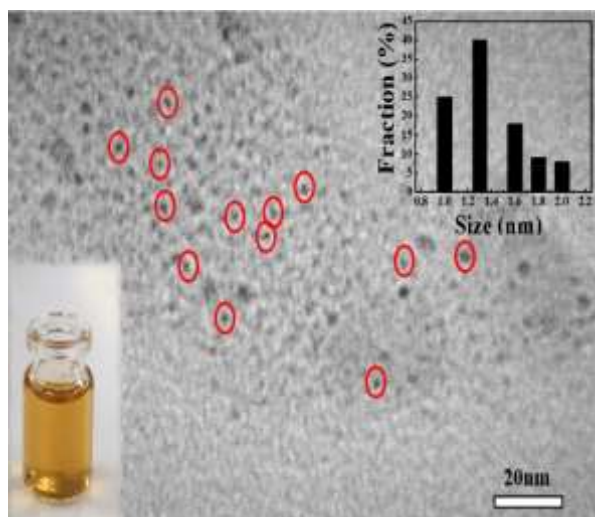


Fig. 2. TEM image and the corresponding size distribution curves of GOQDs.

## III. RESULTS AND DISCUSSION

In our paper, the GOQDs are prepared using CA powder by 5 min of heating at 200°C. The TEM image and the corresponding size distribution plot of the GOQDs for sample with pH=10 are shown in Fig. 2. Results reveal that QDs have quasi-spherical shapes with a nearly monodisperse size distribution. The mean diameters of the GOQDs is about 2 nm. Due to the presence of  $-\text{COOH}$  groups at the edges thereof, the particles are well dispersed, which can be seen in the Fig. 2.

The synthesized GOQDs were characterized by FTIR analysis. Fig. 3 shows the FTIR spectrum of QDs in the range of 400–4000  $\text{cm}^{-1}$  for pH=10. The FTIR spectrum of GOQDs shows the presence of C–O ( $\nu\text{C}-\text{O}$  at 1048, and 1096  $\text{cm}^{-1}$ ), C=O in carboxylic acid, C–OH ( $\nu\text{C}-\text{OH}$  at 1399  $\text{cm}^{-1}$ ), and carbonyl moieties ( $\nu\text{C}=\text{O}$  at 1565 and 1737  $\text{cm}^{-1}$ ), band around 2972  $\text{cm}^{-1}$  due to both  $\text{sp}^2$  and  $\text{sp}^3$  C–H modes, and the O–H band (broadband around 3420  $\text{cm}^{-1}$ ) [29]. According to the articles, pH does not have an effect on reducing and increasing the grafts and functional groups of GOQDs and may only cause shifts [30].

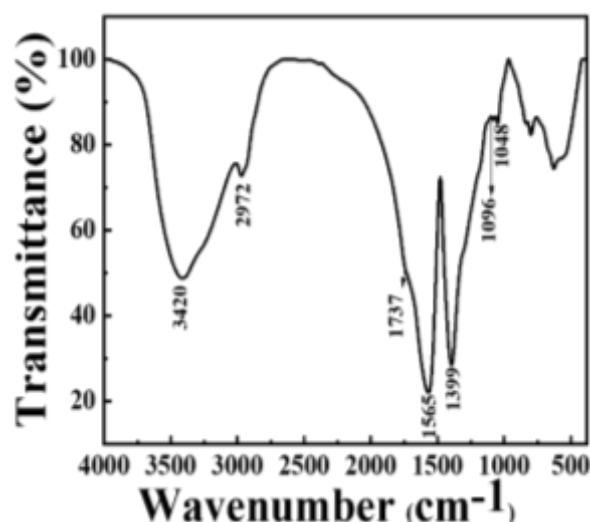


Fig. 3. FTIR spectrum of GOQDs.

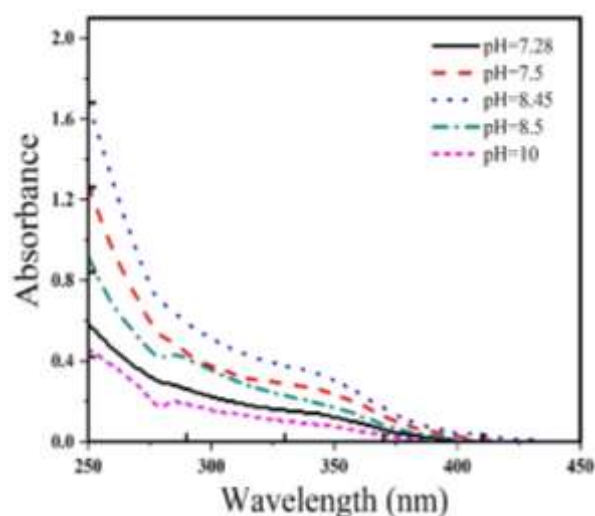


Fig. 4. Absorption spectra of GOQDs in the various pH values.



Figure 4 shows the absorption spectra of GOQDs with various pH values in the aqueous solution. For GOQDs in the Fig. 4, as can be seen, GOQDs show the broad absorption in the UV region with a tail that extends to the visible range. The GOQDs show two shoulder peaks at 210-230 and 340 nm. The absorption peaks at 210-230 nm can be ascribed to  $\pi-\pi^*$

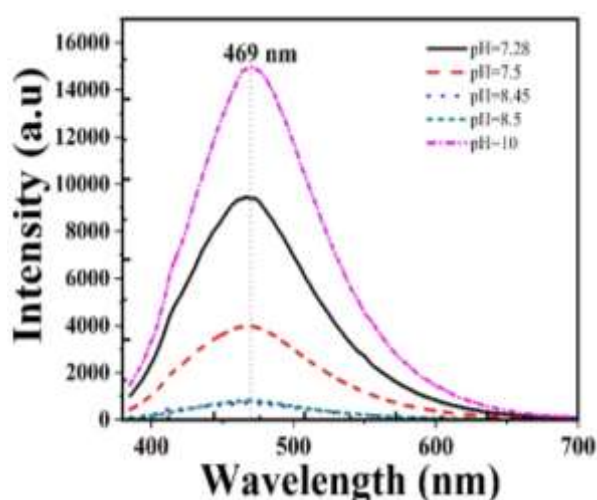


Fig. 5. Fluorescence spectra of GOQDs in the various pH values.

Electronic transitions of the aromatic C–C bonds in their structure. The absorption shoulder peaks at 340 nm can be related to the functional groups on their edges, correctly,  $\pi^*-n$  transitions of the C=O bonds [29]. The GOQDs show a peak in the range of 286 to 300 nm, due to the absorption of graphite structure in water, similar to GO. The absorption intensity in this range for different pHs is increasing and decreasing. Increased adsorption indicates that the electronic connections will be restored after rupture [31]. The UV–Vis spectra proved that on increasing the pH to 8.5, the absorption peaks become broader which caused the increase of distribution of particle size along with the pH. After this pH, the intensity of the peaks decreases up to 10.

Figure 5 presents the room temperature PL spectra of as-prepared GOQDs with various pH values. PL spectra are recorded with an excitation wavelength of 360 nm. The PL spectra of the most carbon materials depends on the wavelength of their excitation. More

and more cases have emerged with excited wavelength that are independent of emission position and can be attributed to their surface chemistry and uniform size. For the all five samples, the emission peaks do not shift in the various pH values, but its intensity decreases by increasing pH from 7.3 to 8.5 and increases with a pH of 10. The emission peak position of all samples was located at 469 nm and the highest photoluminescence intensity related to GOQDs with pH=10. This immutability in the peak position of the PL at different pHs indicates the independence of GOQDs relative to the pH. The primary source of the emission peak position of these materials may be free zigzag modes with a triple carbon state [32]. By increasing the pH, the increasing number of hydroxyl groups gets adhered to the QDs and resulting in an increase in the size of the QDs and lower PL. But, it may be attributed to the fact that in the case of pH=10, the presence of hydroxyl groups impeded the aggregation of the carbon nanoparticles and caused an increase in the size of the nanoparticles and PL. Thus, the pH of the CA solution played a vital role in the synthesis of GOQDs and their optical properties.

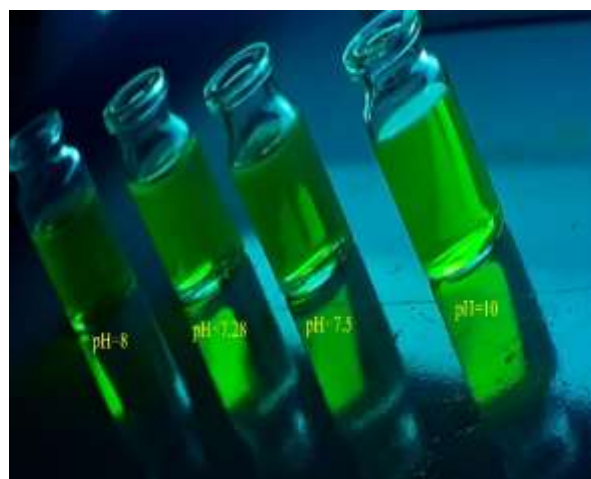


Fig. 6. Photographs of GOQDs under UV light ( $\lambda = 360\text{nm}$ ) in various pH values.

Figure 6 indicates that the solutions of GOQDs are transparent under UV irradiation, and large particles are not observed in them. Therefore, these QDs are well-dispersed in the aqueous phase. The as-prepared GOQDs solutions emit bright green color under UV illumination. The

emission colors of these GOQDs are consistent with the corresponding PL spectra in the Fig. 5. Also, the emission color of all samples is unchanged in various pH values that compatible with the results of PL spectra in Fig. 5. The high transparency of the solutions indicates that the size of the QDs is less than 10 nm.

#### IV. CONCLUSION

GOQDs were synthesized in the various pH values by pyrolyzing CA. Then, the effect of pH on physical and optical properties of synthesized QDs was investigated by FTIR, TEM, and optical spectroscopy. The FTIR analysis of the samples did not change with pH. TEM analysis showed that the particles size diameter was almost the same and less than 5 nm. The UV-Vis spectra of all samples showed two absorption peaks,  $n-\pi^*$  and  $\pi-\pi^*$  transitions, related to C = O and C = C bonds, which are observed in different pHs. The PL spectra of samples showed that the PL peak position did not change in the various pH values, indicating the independence of the PL of the GOQDs from the pH.

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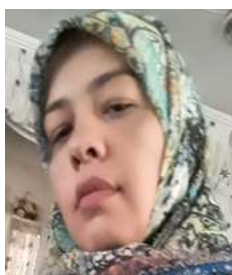
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